

LAMPIRAN

Lampiran 1 Data Perhitungan

1. Perhitungan Pembuatan Larutan Standar

1.1 Konsentrasi larutan induk

Diketahui:

- Massa = 20 mg
- Volume = 100 ml = 0,1 L

Ditanya: konsentrasi

Jawab :

$$1 \text{ ppm} = 1 \text{ mg/L}$$

$$\text{ppm} = \frac{20 \text{ mg}}{0,1 \text{ L}}$$

$$\text{ppm} = 200 \text{ ppm}$$

Jadi, konsentrasi larutan induk untuk membuat deret larutan standar sebesar 200 ppm.

1.2 Konsentrasi Deret Larutan Standar

a) Pembuatan larutan standar 1 ppm

Diketahui :

- $V_1 = 0,05 \text{ ml}$
- $M_1 = 200 \text{ ppm}$
- $V_2 = 10 \text{ ml}$

Ditanya : konsentrasi

Jawab :

$$M_2 = \frac{M_1 \times V_1}{V_2}$$

$$M_2 = \frac{200 \text{ ppm} \times 0,05 \text{ ml}}{10 \text{ ml}}$$

$$\text{ppm} = 1 \text{ ppm}$$

b) Pembuatan larutan standar 3 ppm

Diketahui :

- $V_1 = 0,15 \text{ ml}$
- $M_1 = 200 \text{ ppm}$
- $V_2 = 10 \text{ ml}$

Ditanya : konsentrasi

Jawab :

$$M_2 = \frac{M_1 \times V_1}{V_2}$$

$$M_2 = \frac{200 \text{ ppm} \times 0,15 \text{ ml}}{10 \text{ ml}}$$

$$\text{ppm} = 3 \text{ ppm}$$

c) Pembuatan larutan standar 5 ppm

Diketahui :

- $V_1 = 0,25 \text{ ml}$
- $M_1 = 200 \text{ ppm}$
- $V_2 = 10 \text{ ml}$

Ditanya : Konsentrasi

Jawab :

$$M_2 = \frac{M_1 \times V_1}{V_2}$$

$$M_2 = \frac{200 \text{ ppm} \times 0,25 \text{ ml}}{10 \text{ ml}}$$

$$\text{ppm} = 5 \text{ ppm}$$

d) Pembuatan larutan standar 7 ppm

Diketahui :

- $V_1 = 0,35 \text{ ml}$
- $M_1 = 200 \text{ ppm}$
- $V_2 = 10 \text{ ml}$

Ditanya : Konsentrasi

Jawab :

$$M_2 = \frac{M_1 \times V_1}{V_2}$$

$$M_2 = \frac{200 \text{ ppm} \times 0,35 \text{ ml}}{10 \text{ ml}}$$

$$\text{ppm} = 7 \text{ ppm}$$

e) Pembuatan larutan standar 9 ppm

Diketahui :

- $V_1 = 0,45 \text{ ml}$
- $M_1 = 200 \text{ ppm}$
- $V_2 = 10 \text{ ml}$

Ditanya : Konsentrasi

Jawab :

$$M_2 = \frac{M_1 \times V_1}{V_2}$$

$$M_2 = \frac{200 \text{ ppm} \times 0,45 \text{ ml}}{10 \text{ ml}}$$

$$\text{ppm} = 9 \text{ ppm}$$

2. Perhitungan Konsentrasi Sampel

a) Penimbangan (kode sampel A1)

Diketahui:

- $m = 1000 \text{ mg}$
- $V = 25 \text{ ml} = 0,025 \text{ L}$

Ditanya : Konsentrasi

Jawab :

$$1 \text{ ppm} = 1 \text{ mg/L}$$

$$\frac{1000 \text{ mg}}{0,025 \text{ L}} = 40.000 \text{ ppm}$$

Diambil 1 mL dan diencerkan 25 mL

$$V_1 \times M_1 = V_2 \times M_2$$

$$1 \text{ ml} \times 40.000 \text{ ppm} = 25 \text{ ml} \times M_2$$

$$M_2 = 1.600 \text{ ppm}$$

b) Penimbangan (kode sampel A2)

Diketahui:

- $m = 1000 \text{ mg}$
- $V = 25 \text{ ml} = 0,025 \text{ L}$

Ditanya : Konsentrasi

Jawab :

$$1 \text{ ppm} = 1 \text{ mg/L}$$

$$\frac{1000 \text{ mg}}{0,025 \text{ L}} = 40.000 \text{ ppm}$$

Diambil 1 mL dan diencerkan 25 mL

$$V_1 \times M_1 = V_2 \times M_2$$

$$1 \text{ ml} \times 40.000 \text{ ppm} = 25 \text{ ml} \times M_2$$

$$M_2 = 1.600 \text{ ppm}$$

c) Penimbangan (kode sampel B1)

Diketahui:

- $m = 1001 \text{ mg}$
- $V = 25 \text{ ml} = 0,025 \text{ L}$

Ditanya : Konsentrasi

Jawab :

$$1 \text{ ppm} = 1 \text{ mg/L}$$

$$\frac{1001 \text{ mg}}{0,025 \text{ L}} = 40.040 \text{ ppm}$$

Diambil 1 mL dan diencerkan 25 mL

$$V_1 \times M_1 = V_2 \times M_2$$

$$1 \text{ ml} \times 40.040 \text{ ppm} = 25 \text{ ml} \times M_2$$

$$M_2 = 1.601,6 \text{ ppm}$$

d) Penimbangan (kode sampel B2)

Diketahui:

- $m = 1001 \text{ mg}$
- $V = 25 \text{ ml} = 0,025 \text{ L}$

Ditanya : Konsentrasi

Jawab :

$$1 \text{ ppm} = 1 \text{ mg/L}$$

$$\frac{1001 \text{ mg}}{0,025 \text{ L}} = 40.040 \text{ ppm}$$

Diambil 1 mL dan diencerkan 25 mL

$$\begin{aligned} V_1 \times M_1 &= V_2 \times M_2 \\ 1 \text{ ml} \times 40.040 \text{ ppm} &= 25 \text{ ml} \times M_2 \\ M_2 &= 1.601,6 \text{ ppm} \end{aligned}$$

e) Penimbangan (kode sampel C1)

Diketahui:

- $m = 1001 \text{ mg}$
- $V = 25 \text{ ml} = 0,025 \text{ L}$

Ditanya : Konsentrasi

Jawab :

$$\begin{aligned} 1 \text{ ppm} &= 1 \text{ mg/L} \\ \frac{1001 \text{ mg}}{0,025 \text{ L}} &= 40.040 \text{ ppm} \end{aligned}$$

Diambil 1 mL dan diencerkan 25 mL

$$\begin{aligned} V_1 \times M_1 &= V_2 \times M_2 \\ 1 \text{ ml} \times 40.040 \text{ ppm} &= 25 \text{ ml} \times M_2 \\ M_2 &= 1.601,6 \text{ ppm} \end{aligned}$$

f) Penimbangan (kode sampel C2)

Diketahui:

- $m = 1001 \text{ mg}$
- $V = 25 \text{ ml} = 0,025 \text{ L}$

Ditanya : Konsentrasi

Jawab :

$$\begin{aligned} 1 \text{ ppm} &= 1 \text{ mg/L} \\ \frac{1001 \text{ mg}}{0,025 \text{ L}} &= 40.040 \text{ ppm} \end{aligned}$$

Diambil 1 mL dan diencerkan 25 mL

$$\begin{aligned} V_1 \times M_1 &= V_2 \times M_2 \\ 1 \text{ ml} \times 40.040 \text{ ppm} &= 25 \text{ ml} \times M_2 \\ M_2 &= 1.601,6 \text{ ppm} \end{aligned}$$

g) Penimbangan (kode sampel D1)

Diketahui:

- $m = 1001 \text{ mg}$
- $V = 25 \text{ ml} = 0,025 \text{ L}$

Ditanya : Konsentrasi

Jawab :

$$\begin{aligned} 1 \text{ ppm} &= 1 \text{ mg/L} \\ \frac{1001 \text{ mg}}{0,025 \text{ L}} &= 40.040 \text{ ppm} \end{aligned}$$

Diambil 1 mL dan diencerkan 25 mL

$$\begin{aligned}V_1 \times M_1 &= V_2 \times M_2 \\1\text{ml} \times 40.040 \text{ ppm} &= 25 \text{ ml} \times M_2 \\M_2 &= 1.601,6 \text{ ppm}\end{aligned}$$

h) Penimbangan (kode sampel D2)

Diketahui:

- $m = 1000 \text{ mg}$
- $V = 25 \text{ ml} = 0,025 \text{ L}$

Ditanya : Konsentrasi

Jawab :

$$\begin{aligned}1 \text{ ppm} &= 1 \text{ mg/L} \\ \frac{1000 \text{ mg}}{0,025 \text{ L}} &= 40.000 \text{ ppm}\end{aligned}$$

Diambil 1 mL dan diencerkan 25 mL

$$\begin{aligned}V_1 \times M_1 &= V_2 \times M_2 \\1\text{ml} \times 40.000 \text{ ppm} &= 25 \text{ ml} \times M_2 \\M_2 &= 1.600 \text{ ppm}\end{aligned}$$

i) Penimbangan (kode sampel E1)

Diketahui:

- $m = 1002 \text{ mg}$
- $V = 25 \text{ ml} = 0,025 \text{ L}$

Ditanya : Konsentrasi

Jawab :

$$\begin{aligned}1 \text{ ppm} &= 1 \text{ mg/L} \\ \frac{1002 \text{ mg}}{0,025 \text{ L}} &= 40.080 \text{ ppm}\end{aligned}$$

Diambil 1 mL dan diencerkan 25 mL

$$\begin{aligned}V_1 \times M_1 &= V_2 \times M_2 \\1\text{ml} \times 40.080 \text{ ppm} &= 25 \text{ ml} \times M_2 \\M_2 &= 1.603,2 \text{ ppm}\end{aligned}$$

j) Penimbangan (kode sampel E2)

Diketahui:

- $m = 1000 \text{ mg}$
- $V = 25 \text{ ml} = 0,025 \text{ L}$

Ditanya : Konsentrasi

Jawab :

$$\begin{aligned}1 \text{ ppm} &= 1 \text{ mg/L} \\ \frac{1000 \text{ mg}}{0,025 \text{ L}} &= 40.000 \text{ ppm}\end{aligned}$$

Diambil 1 mL dan diencerkan 25 mL

$$V_1 \times M_1 = V_2 \times M_2$$

$$1 \text{ ml} \times 40.000 \text{ ppm} = 25 \text{ ml} \times M_2$$

$$M_2 = 1.600 \text{ ppm}$$

k) Penimbangan (kode sampel F1)

Diketahui:

- $m = 1003 \text{ mg}$
- $V = 25 \text{ ml} = 0,025 \text{ L}$

Ditanya : Konsentrasi

Jawab :

$$1 \text{ ppm} = 1 \text{ mg/L}$$

$$\frac{1003 \text{ mg}}{0,025 \text{ L}} = 40.120 \text{ ppm}$$

Diambil 1 mL dan diencerkan 25 mL

$$V_1 \times M_1 = V_2 \times M_2$$

$$1 \text{ ml} \times 40.120 \text{ ppm} = 25 \text{ ml} \times M_2$$

$$M_2 = 1.604,8 \text{ ppm}$$

l) Penimbangan (kode sampel F2)

Diketahui:

- $m = 1004 \text{ mg}$
- $V = 25 \text{ ml} = 0,025 \text{ L}$

Ditanya : Konsentrasi

Jawab :

$$1 \text{ ppm} = 1 \text{ mg/L}$$

$$\frac{1004 \text{ mg}}{0,025 \text{ L}} = 40.160 \text{ ppm}$$

Diambil 1 mL dan diencerkan 25 mL

$$V_1 \times M_1 = V_2 \times M_2$$

$$1 \text{ ml} \times 40.160 \text{ ppm} = 25 \text{ ml} \times M_2$$

$$M_2 = 1.606,4 \text{ ppm}$$

3) Nilai Absorbansi Pada Sampel

Diketahui :

- nilai absorbansi

- A1 = 0,216
- A2 = 0,210
- B1 = 0,518
- B2 = 0,465
- C1 = -0,000
- C2 = 0,013
- D1 = -0,693
- D2 = -0,694
- E1 = 0,210

- E2 = 0,277
- F1 = -0,714
- F2 = -0,696

3.1 Perhitungan Konsentrasi Sebenarnya (ppm)

- Persamaan kurva kalibrasi = $y = 0,101222x + 0,0259018$

Ditanya : konsentrasi sebenarnya pada sampel?

a) A1

Jawab :

$$y = 0,101222x + 0,0259018$$

$$0,216 = 0,101222x + 0,0259018$$

$$0,101222x = 0,2419018$$

$$x = \frac{0,2419018}{0,101222}$$

$$x = 2,3898 \text{ ppm}$$

b) A2

Jawab :

$$y = 0,101222x + 0,0259018$$

$$0,210 = 0,101222x + 0,0259018$$

$$0,101222x = 0,2359018$$

$$x = \frac{0,2359018}{0,101222}$$

$$x = 2,3305 \text{ ppm}$$

c) B1

Jawab :

$$y = 0,101222x + 0,0259018$$

$$0,518 = 0,101222x + 0,0259018$$

$$0,101222x = 0,5439018$$

$$x = \frac{0,5439018}{0,101222}$$

$$x = 5,3733 \text{ ppm}$$

d) B2

Jawab :

$$y = 0,101222x + 0,0259018$$

$$0,465 = 0,101222x + 0,0259018$$

$$0,101222x = 0,4909018$$

$$x = \frac{0,4909018}{0,101222}$$

$$x = 4,8497 \text{ ppm}$$

e) C1

Jawab :

$$y = 0,101222x + 0,0259018$$

$$0 = 0,101222x + 0,0259018$$

$$0,101222x = 0,0259018$$

$$x = \frac{0,0259018}{0,101222}$$

$$x = 0,2558 \text{ ppm}$$

f) C2

Jawab :

$$y = 0,101222x + 0,0259018$$

$$0,013 = 0,101222x + 0,0259018$$

$$0,101222x = 0,0389018$$

$$x = \frac{0,0389018}{0,101222}$$

$$x = 0,3843 \text{ ppm}$$

g) D1

Jawab :

$$y = 0,101222x + 0,0259018$$

$$-0,693 = 0,101222x + 0,0259018$$

$$-0,101222x = 0,7189018$$

$$x = \frac{0,7189018}{-0,101222}$$

$$x = -7,1022 \text{ ppm}$$

h) D2

Jawab :

$$y = 0,101222x + 0,0259018$$

$$-0,694 = 0,101222x + 0,0259018$$

$$-0,101222x = 0,7199018$$

$$x = \frac{0,7199018}{-0,101222}$$

$$x = -7,1121 \text{ ppm}$$

i) E1

Jawab :

$$y = 0,101222x + 0,0259018$$

$$0,210 = 0,101222x + 0,0259018$$

$$0,101222x = 0,2359018$$

$$x = \frac{0,2359018}{0,101222}$$

$$x = 2,3305 \text{ ppm}$$

j) E2

Jawab :

$$y = 0,101222x + 0,0259018$$

$$0,277 = 0,101222x + 0,0259018$$

$$0,101222x = 0,3029018$$

$$x = \frac{0,3029018}{0,101222}$$

$$x = 2,9924 \text{ ppm}$$

k) F1

Jawab :

$$y = 0,101222x + 0,0259018$$

$$-0,714 = 0,101222x + 0,0259018$$

$$-0,101222x = 0,7399018$$

$$x = \frac{0,7399018}{-0,101222}$$

$$x = -7,3096 \text{ ppm}$$

l) F2

Jawab :

$$y = 0,101222x + 0,0259018$$

$$-0,696 = 0,101222x + 0,0259018$$

$$-0,101222x = 0,7219018$$

$$x = \frac{0,7219018}{-0,101222}$$

$$x = -7,1318 \text{ ppm}$$

4. Perhitungan Konsentrasi Sebenarnya (ppm) Pada Standard

- Persamaan kurva kalibrasi = $y = 0,101222x + 0,0259018$

Ditanya : konsentrasi sebenarnya pada standard?

a) 1 ppm

Jawab :

$$y = 0,101222x + 0,0259018$$

$$0,132 = 0,101222x + 0,0259018$$

$$0,101222x = 0,1579018$$

$$x = \frac{0,1579018}{0,101222}$$

$$x = 1,5599 \text{ ppm}$$

b) 3 ppm

Jawab :

$$y = 0,101222x + 0,0259018$$

$$0,306 = 0,101222x + 0,0259018$$

$$0,101222x = 0,3319018$$

$$x = \frac{0,3319018}{0,101222}$$

$$x = 3,2789 \text{ ppm}$$

c) 5 ppm

Jawab :

$$y = 0,101222x + 0,0259018$$

$$0,542 = 0,101222x + 0,0259018$$

$$0,101222x = 0,5679018$$

$$x = \frac{0,5679018}{0,101222}$$

$$x = 5,6104 \text{ ppm}$$

d) 7 ppm

Jawab :

$$y = 0,101222x + 0,0259018$$

$$0,763 = 0,101222x + 0,0259018$$

$$0,101222x = 0,7889018$$

$$x = \frac{0,7889018}{0,101222}$$

$$x = 7,7933 \text{ ppm}$$

e) 9 ppm

Jawab :

$$y = 0,101222x + 0,0259018$$

$$0,916 = 0,101222x + 0,0259018$$

$$0,101222x = 0,9419018$$

$$x = \frac{0,9419018}{0,101222}$$

$$x = 9,3053 \text{ ppm}$$

5. Perhitungan Kadar Sampel (%)

a) Kadar sampel A1

Diketahui :

- Konsentrasi perhitungan = 1.600 ppm
- Konsentrasi sebenarnya = 2,3898 ppm

Ditanya : kadar sampel (%)

Jawab :

$$\begin{aligned} \% \text{ sampel} &= \frac{\text{konsentrasi sebenarnya}}{\text{konsentrasi perhitungan}} \times 100 \\ &= \frac{2,3898 \text{ ppm}}{1.600 \text{ ppm}} \times 100 \\ &= 0,149 \% \end{aligned}$$

b) Kadar sampel A2

Diketahui :

- Konsentrasi perhitungan = 1.600 ppm
- Konsentrasi sebenarnya = 2,3305 ppm

Ditanya : kadar sampel (%)

Jawab :

$$\begin{aligned} \% \text{ sampel} &= \frac{\text{konsentrasi sebenarnya}}{\text{konsentrasi perhitungan}} \times 100 \\ &= \frac{2,3305 \text{ ppm}}{1.600 \text{ ppm}} \times 100 \end{aligned}$$

$$= 0,145 \%$$

c) Kadar sampel B1

Diketahui :

- Konsentrasi perhitungan = 1.601,6 ppm

- Konsentrasi sebenarnya = 5,3733 ppm

Ditanya : kadar sampel (%)

Jawab :

$$\begin{aligned} \% \text{ sampel} &= \frac{\text{konsentrasi sebenarnya}}{\text{konsentrasi perhitungan}} \times 100 \\ &= \frac{5,3733 \text{ ppm}}{1.601,6 \text{ ppm}} \times 100 \\ &= 0,335 \% \end{aligned}$$

d) Kadar sampel B2

Diketahui :

- Konsentrasi perhitungan = 1.601,6 ppm

- Konsentrasi sebenarnya = 4,8497 ppm

Ditanya : kadar sampel (%)

Jawab :

$$\begin{aligned} \% \text{ sampel} &= \frac{\text{konsentrasi sebenarnya}}{\text{konsentrasi perhitungan}} \times 100 \\ &= \frac{4,8497 \text{ ppm}}{1.601,6 \text{ ppm}} \times 100 \\ &= 0,302 \% \end{aligned}$$

e) Kadar sampel C1

Diketahui :

- Konsentrasi perhitungan = 1.601,6 ppm

- Konsentrasi sebenarnya = 0,2558 ppm

Ditanya : kadar sampel (%)

Jawab :

$$\begin{aligned} \% \text{ sampel} &= \frac{\text{konsentrasi sebenarnya}}{\text{konsentrasi perhitungan}} \times 100 \\ &= \frac{0,2558 \text{ ppm}}{1.601,6 \text{ ppm}} \times 100 \\ &= 0,015 \% \end{aligned}$$

f) Kadar sampel C2

Diketahui :

- Konsentrasi perhitungan = 1.601,6 ppm

- Konsentrasi sebenarnya = 0,3843 ppm

Ditanya : kadar sampel (%)

Jawab :

$$\% \text{ sampel} = \frac{\text{konsentrasi sebenarnya}}{\text{konsentrasi perhitungan}} \times 100$$

$$= \frac{0,3843 \text{ ppm}}{1.601,6 \text{ ppm}} \times 100$$

$$= 0,023 \%$$

g) Kadar sampel D1

Diketahui :

- Konsentrasi perhitungan = 1.601,6 ppm
- Konsentrasi sebenarnya = -7,1022 ppm

Ditanya : kadar sampel (%)

Jawab :

$$\% \text{ sampel} = \frac{\text{konsentrasi sebenarnya}}{\text{konsentrasi perhitungan}} \times 100$$

$$= \frac{-7,1022 \text{ ppm}}{1.601,6 \text{ ppm}} \times 100$$

$$= -0,443 \%$$

h) Kadar sampel D2

Diketahui :

- Konsentrasi perhitungan = 1.600 ppm
- Konsentrasi sebenarnya = -7,1121 ppm

Ditanya : kadar sampel (%)

Jawab :

$$\% \text{ sampel} = \frac{\text{konsentrasi sebenarnya}}{\text{konsentrasi perhitungan}} \times 100$$

$$= \frac{-7,1121 \text{ ppm}}{1.600 \text{ ppm}} \times 100$$

$$= -0,444 \%$$

i) Kadar sampel E1

Diketahui :

- Konsentrasi perhitungan = 1.603,2 ppm
- Konsentrasi sebenarnya = 2,3305 ppm

Ditanya : kadar sampel (%)

Jawab :

$$\% \text{ sampel} = \frac{\text{konsentrasi sebenarnya}}{\text{konsentrasi perhitungan}} \times 100$$

$$= \frac{2,3305 \text{ ppm}}{1.603,2 \text{ ppm}} \times 100$$

$$= 0,145 \%$$

j) Kadar sampel E2

Diketahui :

- Konsentrasi perhitungan = 1.600 ppm
- Konsentrasi sebenarnya = 2,9924 ppm

Ditanya : kadar sampel (%)

Jawab :

$$\% \text{ sampel} = \frac{\text{konsentrasi sebenarnya}}{\text{konsentrasi perhitungan}} \times 100$$

$$= \frac{2,9924 \text{ ppm}}{1.600 \text{ ppm}} \times 100$$

$$= 0,187 \%$$

k) Kadar sampel F1

Diketahui :

- Konsentrasi perhitungan = 1.604,8 ppm
- Konsentrasi sebenarnya = -7,3096 ppm

Ditanya : kadar sampel (%)

Jawab :

$$\% \text{ sampel} = \frac{\text{konsentrasi sebenarnya}}{\text{konsentrasi perhitungan}} \times 100$$

$$= \frac{-7,3096 \text{ ppm}}{1.604,8 \text{ ppm}} \times 100$$

$$= -0,455 \%$$

l) Kadar sampel F2

Diketahui :

- Konsentrasi perhitungan = 1.606,4 ppm
- Konsentrasi sebenarnya = -7,1318 ppm

Ditanya : kadar sampel (%)


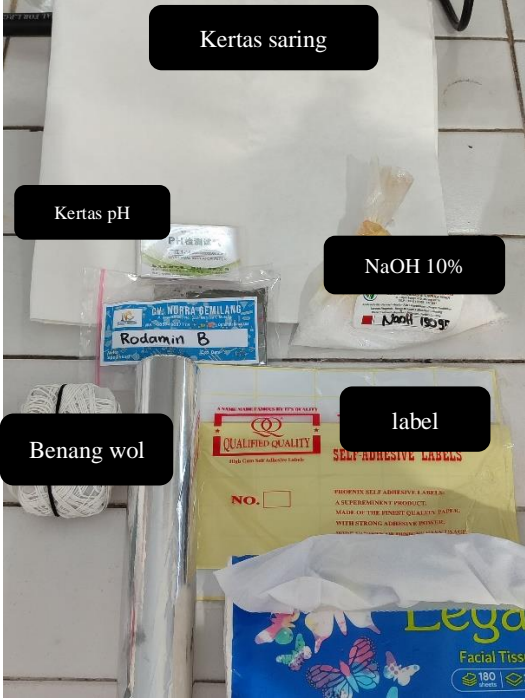
Jawab :

$$\% \text{ sampel} = \frac{\text{konsentrasi sebenarnya}}{\text{konsentrasi perhitungan}} \times 100$$











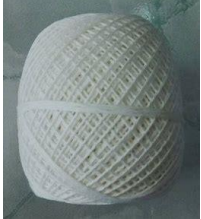




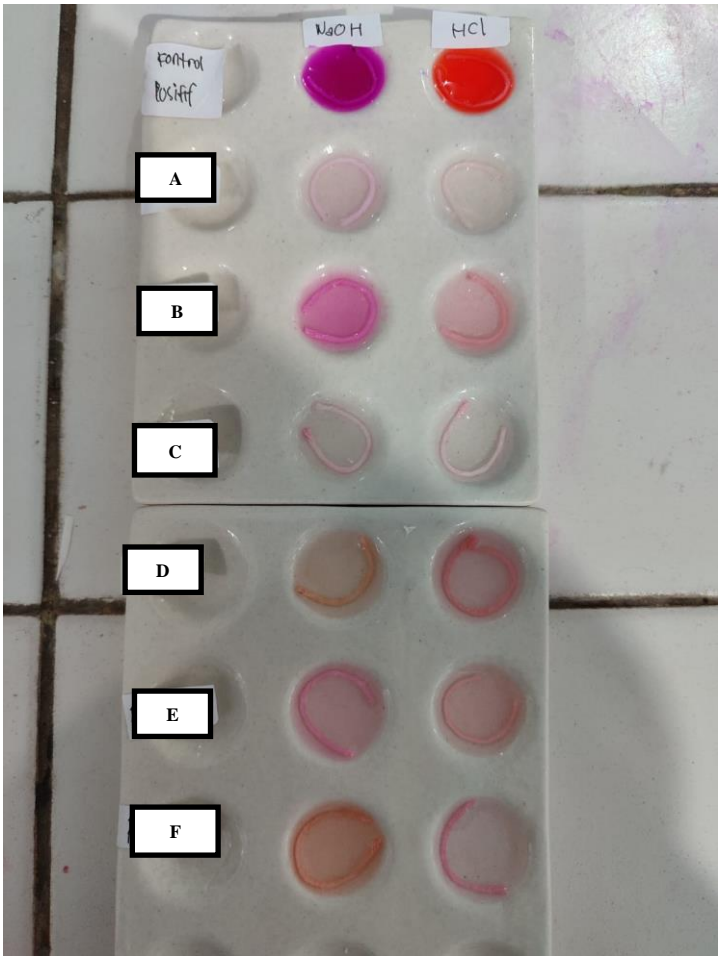
$$= \frac{-7,1318 \text{ ppm}}{1.606,4 \text{ ppm}} \times 100$$

$$= -0,443 \%$$




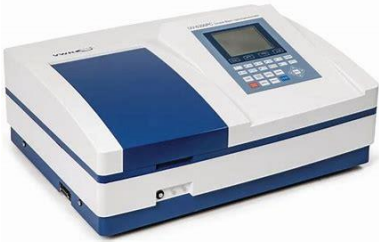



Lampiran 2 Gambar Bahan

Bahan	Sampel
	
	<p>Keterangan:</p> <p>Sampel dengan kode A, B, C = Non BPOM</p> <p>Sampel dengan kode D, E, F = Ber-BPOM</p>

Lampiran 3 Uji Pewarnaan

Alat	Hasil Pengerjaan
 <p>Corong</p>	
 <p>Plat tetes</p>	
 <p>Pipet ukur</p>	
 <p>Pipet tetes</p>	
 <p>Batang pengaduk</p>	
 <p>Benang wol</p>	
 <p>Hot plate</p>	
 <p>Beaker glass 100 mL</p>	

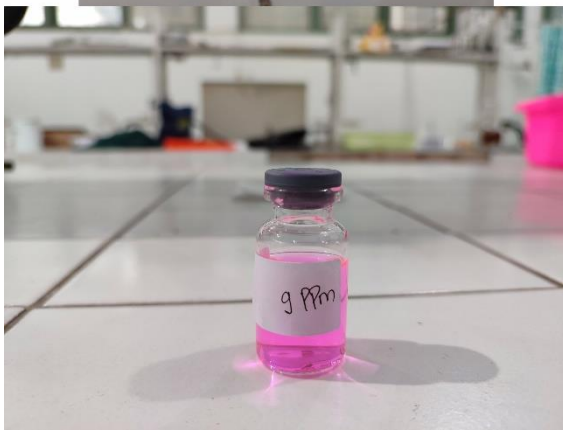
Lampiran 4 Gambar Alat Pengujian Spektrofotometri UV-Vis

Alat		
		
Kuvet	Mikropipet	Vial 10 mL
		
Spektrofotometri UV-Vis	Labu takar 100 mL	Batang pengaduk
		
Labu takar 10 mL		

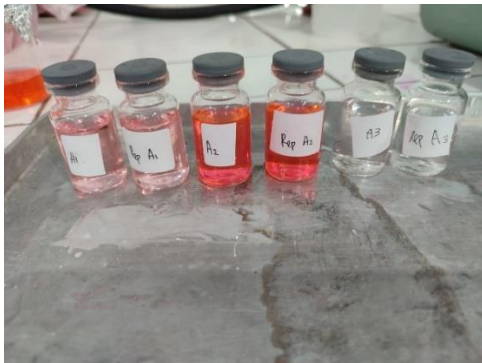
Lampiran 5 Uji Spektrofotometri UV-Vis

Hasil pengujian

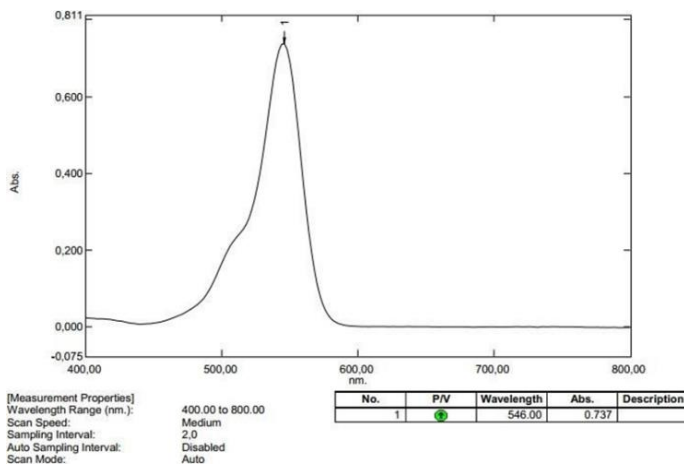
Deret larutan baku Rhodamin B



Larutan sampel akan diuji pada Spektrofotometri UV-Vis



Panjang gelombang maksimum



Hasil pengukuran absorbansi larutan baku

$$y = 0,101222 x + 0,0259018$$

$$r^2 = 0,99529$$

Standard Table

	Sample ID	Type	Ex	Conc	WL546,0	Wgt.Factor	Comments
1	1	Standard		1.000	0.132	1.000	
2	3	Standard		3.000	0.306	1.000	
3	5	Standard		5.000	0.542	1.000	
4	7	Standard		7.000	0.763	1.000	
5	9	Standard		9.000	0.916	1.000	
6							

Hasil pengukuran absorbansi larutan sampel

Sequence No.

Sample Table

	Sample ID	Type	Ex	Conc	WL546,0	Comments
1	A1	Unknown		1.876	0.216	
2	Rep1	Unknown		1.816	0.210	
3	A3	Unknown		-0.259	-0.000	
4	Rep3	Unknown		-0.128	0.013	
5	a22	Unknown		4.866	0.518	
6	a22rep	Unknown		4.340	0.465	
7	B1	Unknown		-7.105	-0.693	
8	RepB1	Unknown		-7.108	-0.694	
9	B2	Unknown		1.818	0.210	
10	RepB2	Unknown		2.485	0.277	
11	B3	Unknown		-7.313	-0.714	
12	RepB3	Unknown		-7.129	-0.696	
13						